Roll No. .....

Question Booklet Number

O. M. R. Serial No.

M. Sc. (Second Semester)
(NEP) EXAMINATION, 2025-26
CHEMISTRY
(Physical Chemistry—II)

Paper Code							
B	0	2	0	8	0	3	T

Questions Booklet Series
C

Time : 1:30 Hours] [*Maximum Marks : 75*

Instructions to the Examinee :

1. Do not open the booklet unless you are asked to do so.

2. The booklet contains 100 questions. Examinee is required to answer 75 questions in the OMR Answer-Sheet provided and not in the question booklet. All questions carry equal marks.

3. Examine the Booklet and the OMR Answer-Sheet very carefully before you proceed. Faulty question booklet due to missing or duplicate pages/questions or having any other discrepancy should be got immediately replaced.

परीक्षार्थियों के लिए निर्देश :

1. प्रश्न-पुस्तिका को तब तक न खोलें जब तक आपसे कहा न जाए।

2. प्रश्न-पुस्तिका में 100 प्रश्न हैं। परीक्षार्थी को 75 प्रश्नों को केवल दी गई OMR आन्सर-शीट पर ही हल करना है, प्रश्न-पुस्तिका पर नहीं। सभी प्रश्नों के अंक समान हैं।

3. प्रश्नों के उत्तर अंकित करने से पूर्व प्रश्न-पुस्तिका तथा OMR आन्सर-शीट को सावधानीपूर्वक देख लें। दोषपूर्ण प्रश्न-पुस्तिका जिसमें कुछ भाग छपने से छूट गए हों या प्रश्न एक से अधिक बार छप गए हों या उसमें किसी अन्य प्रकार की कमी हो, तो उसे तुरन्त बदल लें।

(Remaining instructions on the last page) **(शेष निर्देश अन्तिम पृष्ठ पर)**

1. According to the theory of Absolute Reaction Rate, the activated complex is in equilibrium with :

(A) The products only

(B) The reactants only

(C) Reactants and products simultaneously

(D) Neither reactants nor products

2. The graph obtained in polarography is called :

(A) Chromatogram

(B) Spectrogram

(C) Polarogram

(D) Oscillogram

3. The working electrode used in classical polarography is :

(A) Platinum electrode

(B) Glass electrode

(C) Dropping mercury electrode

(D) Silver electrode

4. The diffusion current in polarography is directly proportional to :

(A) Square of concentration

(B) Concentration

(C) Log current

(D) Square root of concentration

5. The Ilkovic equation is used to calculate :

(A) Migration current

(B) Diffusion current

(C) Residual current

(D) Limiting potential

6. The turbidity of a polymer solution measures :

(A) A light absorbed by solution

(B) Light transmitted by the solution

(C) Light scattered by the solution

(D) None of the above

7. Monomer of $\left[\begin{array}{l} CH_3 \\ | \\ -C-CH_2- \\ | \\ CH_3 \end{array} \right]$ is :

(A) 2-methyl propane

(B) Styrene

(C) Propylene

(D) Ethene

8. Which of the following is used in tyre cords ?

(A) Terylene

(B) Polyethylene

(C) Polypropylene

(D) Nylon-6

9. Bakelite is prepared by the reaction between :

(A) Urea and Formaldehyde

(B) Tetramethylene glycol and hexamethylene isocyanate

(C) Phenol and Formaldehyde

(D) Ethylene glycol and dimethyl tetraphthalate

10. The catalyst used in the manufacture of polythene by Zeigler method is :

(A) Lithium tetrachloride and triphenyl aluminium

(B) Titanium tetrachloride and triphenyl aluminium

(C) Titanium oxide

(D) Titanium isoperoxide

11. The weakest interparticle forces are present in :

(A) Thermosetting polymers

(B) Thermoplastic polymers

(C) Fibres

(D) Elastomers

12. If λ_∞ and λ_V are the equivalent conductances at infinite dilution and at V dilution, the degree of dilution, ∞, is given by :

(A) $\alpha = \frac{\lambda_\infty}{\lambda_v}$

(B) $\alpha = \frac{\lambda_\infty}{\lambda_v^2}$

(C) $\alpha = \frac{\lambda_v}{\lambda_\infty}$

(D) None of the above

13. Kohlrausch's Law can be used to determine :

(A) λ_∞ for weak electrolytes

(B) Absolute ionic mobilities

(C) Solubility of sparingly soluble salts

(D) All of the above

14. The Hittorf's Rule can be represented by the expression :

(A) $\frac{\text{Fall around anode}}{\text{Rise around cathode}} = \frac{\text{Speed of cation}}{\text{Speed of anion}}$

(B) $\frac{\text{Fall around anode}}{\text{Fall around cathode}} = \frac{\text{Speed of anion}}{\text{Speed of cation}}$

(C) $\frac{\text{Fall around cathode}}{\text{Fall around anode}} = \frac{\text{Speed of cation}}{\text{Speed of anion}}$

(D) $\frac{\text{Rise around cathode}}{\text{Fall around anode}} = \frac{\text{Speed of anion}}{\text{Speed of cation}}$

15. A 0·01 N solution of a weak monobasic acid dissociates to the extent of 10%. Calculate the dissociation constant of the acid :

(A) $1{\cdot}61 \times 10^{-4}$ moles/L

(B) $2{\cdot}61 \times 10^{+4}$ moles/L

(C) $1{\cdot}11 \times 10^{-4}$ moles/L

(D) $2{\cdot}96 \times 10^{+4}$ moles/L

16. Corrosion occurs due to mixture of iron and :

(A) Sulphur

(B) Hydrogen

(C) Nitrogen

(D) Oxygen

17. Which of the following are the protective measures to prevent the metals from corrosion ?

(A) Non-metallic coatings

(B) Coating with paint

(C) Enamelling

(D) All of the above

18. What kind of corrosion is difficult to monitor and is very dangerous for metals ?

(A) Pitting corrosion

(B) Stress corrosion

(C) Galvanic corrosion

(D) Crevice corrosion

19. An oxide layer is considered as protective if ratio of volume of elementary cell of oxide to volume of elementary cell of base metal :

(A) is less than 1

(B) is greater than 10

(C) is greater than 1

(D) is greater than 3

20. The diffusion current is measured in late life period of mercury drop in :

(A) Pulsed Polarography

(B) Differential Pulsed Polarography

(C) Tast Polarography

(D) Voltammetry

21. ECG (Electrocardiogram) was developed first by :

(A) Wilhelm His

(B) Steward

(C) Hubert Mann

(D) William Einthoven

22. The Fermi-Level in an intrinsic semi conductor is :

(A) Closer to valence band

(B) Nearly midway between conduction and valence band

(C) Closer to conduction band

(D) Within the valence band

23. For an intrinsic semiconductor (n_h and n_e are the number of holes per unit volume and the number of electrons per unit volume respectively) :

(A) $n_h < n_e$

(B) $n_h = n_e$

(C) $n_h = \frac{n_e}{2}$

(D) $n_h > n_e$

24. The diffusion term in Nernst-Planck equation is based on :

(A) Ohm's Law

(B) Raoult's Law

(C) Fick's First Law

(D) Henry's Law

25. When light interacts with a semiconductor-solution interface, which of the following effects is most likely to occur ?

(A) Increased conductivity in semi-conductor due to photo-generated charge carriers

(B) A decrease in band-gap energy of the semiconductor

(C) An increase in the refractive index of solution

(D) A decrease in number of charge carriers in the semiconductor

26. Which special type of diode is capable of both amplification and oscillation ?

(A) Point contact diode

(B) Zener diode

(C) Junction diode

(D) Tunnel diode

27. Which principle describes the behaviour of charge carriers in a semi-conductor ?

(A) Ohm's Law

(B) Fermi-Dirac Statistics

(C) Hooke's Law

(D) Archimedes' Principle

28. Which plot is derived from the Butler-Volmer equation ?

(A) Nernst plot

(B) Arrhenius plot

(C) Tafel plot

(D) Debye-Huckel plot

29. Which one of the following is the disadvantage of ion-implantations over diffusion doping ?

(A) It is a low temperature process

(B) Point imperfection are not produced

(C) Shallow doping is possible

(D) Gattering is possible

30. The standard hydrogen electrode has potential :

(A) 1 V

(B) 0·5 V

(C) 0 V

(D) – 1 V

31. Transport number of an ion is :

(A) Total ions in solution

(B) Fraction of current carried by that ion

(C) Charge on the ion

(D) Resistance of ion

32. A galvanic cell converts :

(A) Electrical energy to chemical energy

(B) Heat energy to electrical energy

(C) Mechanical energy to chemical energy

(D) Chemical energy to electrical energy

33. The equation relating EMF and concentration is :

(A) Arrhenius equation

(B) Henderson equation

(C) Nernst equation

(D) Debye-Huckel equation

34. According to Kohlrausch's Law at infinite dilution :

(A) Ions interact strongly

(B) Conductance decreases

(C) Each ion contributes independently

(D) Electrolyte becomes non-conducting

35. Molar concentration increases with dilution because :

(A) Number of ions increases

(B) Degree of ionization increases

(C) Viscosity increases

(D) Resistance increases

36. Intensity of scattered light depends on :

(A) Molecular size

(B) Molecular weight

(C) Concentration

(D) All of the above

37. Diffusion method of determining mol. wt. is based on :

(A) Laplace equation

(B) BET equation

(C) Arrhenius equation

(D) Stokes–Einstein equation

38. Osmometry is most suitable for polymers of :

(A) Very high mol. wt.

(B) Very low mol. wt.

(C) Moderate mol. wt.

(D) Cross-linked polymers

39. Osmotic pressure measurement gives :

(A) Weight average molecular weight

(B) Number average molecular weight

(C) Viscosity average molecular weight

(D) Z-average molecular weight

40. Tacticity in polymers arises due to the presence of :

(A) Double bonds

(B) Chiral carbon atoms

(C) Ionic bonds

(D) Crosslinking

41. A polymer having all substituent groups on the same site of the polymer chain is :

(A) Atactic

(B) Syndiotactic

(C) Isotactic

(D) Random

42. Which of the following is a configurational isomerism in polymers ?

(A) Cis-trans isomerism

(B) Conformational change

(C) Bond rotation

(D) Chain folding

43. Rubber is an example of :

(A) Thermoplastic

(B) Elastomer

(C) Thermosetting plastic

(D) Fibre

44. Which of the following is a thermosetting polymer ?

(A) PVC

(B) Polystyrene

(C) Polyethylene

(D) Bakelite

45. Example of Addition polymerization is :

(A) Nylon 6, 6

(B) Bakelite

(C) Polyethylene

(D) Dacron

46. Degree of polymerization is :

(A) Number of monomer units in a polymer chain

(B) Molecular weight of monomer

(C) Viscosity of polymer solution

(D) Number of branches in polymer

47. Polymerization in which small molecules like water are eliminated is :

(A) Addition polymerization

(B) Free radical polymerization

(C) Condensation polymerization

(D) Chain transfer

48. Addition of electrolyte to an ionic surfactant solution :

(A) Increases CMC

(B) Decreases CMC

(C) Has no effect

(D) Destroys micelles

49. Below CMC, surfactant molecules exist mainly as :

(A) Micelles

(B) Dimers

(C) Monomers

(D) Colloidal particles

50. Critical Micelle Concentration (CMC) is defined as :

(A) Concentration at which surfactant completely dissolves

(B) Concentration at which micelles start forming

(C) Concentration at which surface tension becomes zero

(D) Maximum solubility of surfactant

51. From BET plot : Slope = 15 and Intercept = 3
Find BET constant (C).
(A) 4
(B) 3
(C) 5
(D) 6

52. BET theory assumes :
(A) Monolayer adsorption only
(B) Multilayer adsorption
(C) No adsorption
(D) Chemical reaction only

53. The BET equation is mainly used to determine :
(A) Molecular weight
(B) Surface area of solids
(C) Viscosity
(D) Surface tension

54. A small bubble compared to a large bubble will have :
(A) Lower internal pressure
(B) Same internal pressure
(C) Higher internal pressure
(D) Zero pressure

55. The unit of surface tension in SI system is :
(A) N/m
(B) N
(C) Pa
(D) J

56. The pressure difference across a curved surface depends on :
(A) Density only
(B) Viscosity only
(C) Surface tensions and radius
(D) Temperature only

57. According to Laplace equation, if the radius of a droplet decreases, the internal pressure :
(A) Decreases
(B) Increases
(C) Remains constant
(D) Becomes zero

58. Zwitterionic surfactants contain :
(A) Only positive charge
(B) Only negative charge
(C) Both positive and negative charges
(D) No charge

59. Hydrophilic-Lipophilic Balance (HLB) is used to :
(A) Measure Viscosity
(B) Classify Surfactants
(C) Determine pH
(D) Measure density

60. Cationic surfactants are commonly used as :
(A) Detergents
(B) Antiseptics
(C) Emulsifiers in food
(D) Sweetening agents

61. The concentration at which micelles begin to form is called :

(A) HLB
(B) Kraft point
(C) Critical Micelle Concentration (CMC)
(D) Cloud point

62. A molecule that has both hydrophilic and hydrophobic parts is called :

(A) Lyophilic
(B) Amphiphilic
(C) Hydrotropic
(D) Electrolytic

63. Surface active agents primarily acts by :

(A) Increasing surface tension
(B) Decreasing surface tension
(C) Increasing viscosity
(D) Decreasing solubility

64. Measuring zeta-potential is useful in determining which property of a liquid formulation ?

(A) Viscosity
(B) Stability
(C) Solubility
(D) Particle size

65. The most commonly used polymeric surfactants used in pharmacy are :

(A) A-B-A block copolymers
(B) B-A-B block copolymers
(C) A-A-A block copolymers
(D) B-B-B block copolymers

66. Which of the following are cationic surfactants ?

(A) Benzethonium chloride
(B) N-dodecyl alanine
(C) Salts of divalent and trivalent metals
(D) All of the above

67. The concentration at which micelle formation occurs is called :

(A) Kraft point
(B) Cloud point
(C) Isoelectric point
(D) CMC

68. Adsorption theory explains which of the following types of catalysis ?

(A) Heterogeneous catalysis
(B) Homogeneous catalysis
(C) Autocatalysis
(D) None of the above

69. In a steady state of a two-dimensional heat flow, if u is independent of t, then the equation reduces to :

(A) Heat flow in three dimensions
(B) Laplace's equation in three dimensions
(C) Laplace's equation in two dimensions
(D) Heat flow in two dimensions

70. Chromatography is a technique based on :

(A) Solubilities of solute

(B) Absorption of solute

(C) Chemical adsorption followed by dispersion

(D) Differential adsorption of different constituents of a mixture

71. What is the major limitation of the Lindemann-Hinshelwood theory ?

(A) It fails to explain high-pressure kinetics

(B) It assumes molecules are rigid spheres

(C) It ignores the internal degrees of freedom and energy dependence of activation

(D) It cannot be used for gases

72. Which reaction demonstrates oscillatory behaviour ?

(A) Hydrogen-bromine

(B) Pyrolysis of acetaldehyde

(C) Belousov-Zhabotinsky

(D) Decomposition of ethane

73. In the Lindemann-Hinshelwood mechanism, the reaction becomes pseudo-first order under which condition ?

(A) Very low pressure

(B) High pressure

(C) Constant volume

(D) Low temperature

74. Homogeneous catalysis refers to a reaction where :

(A) Catalyst and reactants are in same phase

(B) Catalyst and reactants are in different phases

(C) Catalyst is always solid

(D) Reaction occurs only in aqueous medium

75. Which of the following is an example of homogeneous catalysis ?

(A) Hydrogenation of oil using Ni

(B) Decomposition of H_2O_2 by MnO_2

(C) Esterification of acetic acid with ethanol using H_2SO_4

(D) Haber process using Fe

76. The value of Steric factor is generally :

(A) Equal to 1

(B) Greater than 1

(C) Less than or equal to 1

(D) Always zero

77. The Steric factor (P) in a chemical reaction accounts for :

(A) Energy of activation

(B) Orientation of colliding molecules

(C) Temperature dependence

(D) Pressure changes

78. A catalyst increases reaction rate by :

(A) Increasing activation energy

(B) Decreasing activation energy

(C) Increasing threshold energy

(D) Stopping collisions

79. The intermediate formed during a reaction with maximum energy is :

(A) Reactant

(B) Product

(C) Activated complex

(D) Catalyst

80. Increasing temperature increases rate of reaction because :

(A) It decreases activation energy

(B) It increases number of effective collisions

(C) It decreases collisions

(D) It stops molecular motion

81. The minimum energy required for a reaction to occur is called :

(A) Threshold energy

(B) Potential energy

(C) Kinetic energy

(D) Bond energy

82. The steady-state assumption is mainly used in the derivation of :

(A) Lineweaver-Burk equation

(B) Michaelis-Menten equation

(C) Arrhenius equation

(D) Henderson-Hasselbatch equation

83. When substrate concentration is much greater than K_m, the reaction is :

(A) First order

(B) Zero order

(C) Second order

(D) Third order

84. At steady state, which of the following is true ?

(A) Rate of ES formation = Rate of ES breakdown

(B) Substrate concentration equals product concentration

(C) Enzyme concentration equals substrate concentration

(D) Reaction has reached equilibrium

85. For a reaction between two ions of same charge sign, increasing ionic strength will :

(A) Decrease rate constant

(B) Increase rate constant

(C) No effect

(D) Make reaction zero order

86. The kinetic salt effect is explained by which equation ?

(A) Arrhenius equation

(B) Nernst equation

(C) Bronsted Bjerrum equation

(D) Henderson equation

87. Enzyme activity is directly proportional to enzyme concentration when :

(A) Substrate is limiting

(B) Substrate is in excess

(C) pH is low

(D) Temperature is zero

88. The Lineweaver-Burk plot is a graph between :

(A) V *vs* [S]

(B) $\frac{1}{V}$ *vs* $\frac{1}{[S]}$

(C) log V *vs* log [S]

(D) V^2 *vs* [S]

89. When substrate concentration equals K_m, the reaction velocity is :

(A) V_{max}

(B) Zero

(C) Half of V_{max}

(D) Double V_{max}

90. K_m represents :

(A) Maximum velocity of reaction

(B) Substrate concentration at half V_{max}

(C) Enzyme concentration

(D) Rate constant of a reaction

91. The relationship between substrate concentration and reaction velocity in enzyme-catalyzed reaction is described by :

(A) Arrhenius equation

(B) Michaelis-Menten equation

(C) Van't Hoff equation

(D) Henderson-Hasselbalch equation

92. What is the main advantage of using Pulse Radiolysis to study fast reactions ?

(A) It allows for the study of reactions in solid state

(B) It provides a very high degree of control over the reaction conditions

(C) It allows for the study of reactions with very fast reaction rates

(D) It allows for the study of reactions in the gas phase

93. In the stopped flow method for measuring fast reaction, the reactants are :

(A) Mixed rapidly and then the reaction is stopped abruptly

(B) Mixed slowly and then the reaction is allowed to proceed normally

(C) Mixed rapidly and then the reaction is allowed to proceed normally

(D) Mixed slowly and then the reaction is stopped abruptly

94. Which of the following techniques is often used in conjugation with NMR to study very fast reactions by allowing the measurement of spectra at different, specific time intervals after mixing ?

(A) Stopped flow NMR

(B) Solid state NMR

(C) Real time monitoring

(D) Ultrafast 2 DNMR

95. The 'NMR time scale' implies that fast reactions can be studied if the rate of reaction is comparable to :

(A) The time it takes to record a spectrum (minutes)

(B) The spin lattice relaxation time

(C) The chemical shift difference (in Hz) between the exchanging sites

(D) The speed of the Fourier transform

96. The effect of temperature on the rate of reaction is given by :

(A) Van't Hoff equation

(B) Clausius–Clapeyron equation

(C) Kirchhoff's equation

(D) Arrhenius equation

97. For a Lindemann-Hinshelwood mechanism, what will be the order of reaction if the rate of deactivation by collision is much greater than the rate of unimolecular decay ?

(A) Zero order

(B) First order

(C) Second order

(D) Third order

98. In the mechanism of reaction

$$H_2 + Br_2 \rightarrow 2HBr$$

the first step is :

(A) Dissociation of Br_2 into Br^- radicals

(B) Dissociation of H_2 into H^+ radicals

(C) Reaction of H^+ radical with Br_2

(D) Reaction of Br^- radical with H_2

99. Activation energy of a reaction whose rate constant at 27ºC gets doubled for 10ºC rise in temperature is :

(A) 1366·85 cal

(B) 12804·5 cal

(C) 13668·5 cal

(D) 1280·45 cal

100. Units of rate constant in the following rate law is :

$r = k$ [A] [B] if concentration is expressed in moles litre^{-1} and time in seconds :

(A) L mol^{-1} sec^{-1}

(B) mol L^{-1} sec^{-1}

(C) mol sec L^{-1}

(D) L sec mol^{-1}

(Only for Rough Work)

4. Four alternative answers are mentioned for each question as—A, B, C & D in the booklet. The candidate has to choose the correct answer and mark the same in the OMR Answer-Sheet as per the direction :

Example :

Question :

Q. 1 (A) ● (C) (D)
Q. 2 (A) (B) ● (D)
Q. 3 (A) ● (C) (D)

Illegible answers with cutting and over-writing or half filled circle will be cancelled.

5. Each question carries equal marks. Marks will be awarded according to the number of correct answers you have.
6. All answers are to be given on OMR Answer Sheet only. Answers given anywhere other than the place specified in the answer sheet will not be considered valid.
7. Before writing anything on the OMR Answer Sheet, all the instructions given in it should be read carefully.
8. After the completion of the examination candidates should leave the examination hall only after providing their OMR Answer Sheet to the invigilator. Candidate can carry their Question Booklet.
9. There will be no negative marking.
10. Rough work, if any, should be done on the blank pages provided for the purpose in the booklet.
11. To bring and use of log-book, calculator, pager and cellular phone in examination hall is prohibited.
12. In case of any difference found in English and Hindi version of the question, the English version of the question will be held authentic.

Impt. : On opening the question booklet, first check that all the pages of the question booklet are printed properly. If there is any discrepancy in the question Booklet, then after showing it to the invigilator, get another question Booklet of the same series.

4. प्रश्न-पुस्तिका में प्रत्येक प्रश्न के चार सम्भावित उत्तर—A, B, C एवं D हैं। परीक्षार्थी को उन चारों विकल्पों में से सही उत्तर छाँटना है। उत्तर को OMR आन्सर-शीट में सम्बन्धित प्रश्न संख्या में निम्न प्रकार भरना है :

उदाहरण :

प्रश्न :

प्रश्न 1 (A) ● (C) (D)
प्रश्न 2 (A) (B) ● (D)
प्रश्न 3 (A) ● (C) (D)

अपठनीय उत्तर या ऐसे उत्तर जिन्हें काटा या बदला गया है, या गोले में आधा भरकर दिया गया, उन्हें निरस्त कर दिया जाएगा।

5. प्रत्येक प्रश्न के अंक समान हैं। आपके जितने उत्तर सही होंगे, उन्हीं के अनुसार अंक प्रदान किये जायेंगे।
6. सभी उत्तर केवल ओ. एम. आर. उत्तर-पत्रक (OMR Answer Sheet) पर ही दिये जाने हैं। उत्तर-पत्रक में निर्धारित स्थान के अलावा अन्यत्र कहीं पर दिया गया उत्तर मान्य नहीं होगा।
7. ओ. एम. आर. उत्तर-पत्रक (OMR Answer Sheet) पर कुछ भी लिखने से पूर्व उसमें दिये गये सभी अनुदेशों को सावधानीपूर्वक पढ़ लिया जाये।
8. परीक्षा समाप्ति के उपरान्त परीक्षार्थी कक्ष निरीक्षक को अपनी OMR Answer Sheet उपलब्ध कराने के बाद ही परीक्षा कक्ष से प्रस्थान करें। परीक्षार्थी अपने साथ प्रश्न-पुस्तिका ले जा सकते हैं।
9. निगेटिव मार्किंग नहीं है।
10. कोई भी रफ कार्य, प्रश्न-पुस्तिका के अन्त में, रफ-कार्य के लिए दिए खाली पेज पर ही किया जाना चाहिए।
11. परीक्षा-कक्ष में लॉग-बुक, कैलकुलेटर, पेजर तथा सेल्युलर फोन ले जाना तथा उसका उपयोग करना वर्जित है।
12. प्रश्न के हिन्दी एवं अंग्रेजी रूपान्तरण में भिन्नता होने की दशा में प्रश्न का अंग्रेजी रूपान्तरण ही मान्य होगा।

महत्वपूर्ण : प्रश्नपुस्तिका खोलने पर प्रथमतः जाँच कर देख लें कि प्रश्न-पुस्तिका के सभी पृष्ठ भलीभाँति छपे हुए हैं। यदि प्रश्नपुस्तिका में कोई कमी हो, तो कक्षनिरीक्षक को दिखाकर उसी सिरीज की दूसरी प्रश्न-पुस्तिका प्राप्त कर लें।